

A series of CBP-derivatives as host materials for blue phosphorescent organic light-emitting diodes

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We report a series of **CBP**-derivatives with superior thermal and electronic properties for the use as host materials for blue electrophosphorescent organic light emitting diodes. We applied a systematic variation of the substitution pattern in the 2- and 2'-position of the biphenyl unit and the 3- and 6-position of the carbazole moieties. In contrast to the crystalline parent compound **CBP**, all methyl and trifluoromethyl substituted derivatives show amorphous behaviour. Substitution in the 2- and 2'-position of the biphenyl causes a twisting of the phenyl rings. Hence, the degree of conjugation of the molecules is limited which leads to enlarged triplet energies of approximately 2.95 eV compared to 2.58 eV for **CBP**. The methyl substitution at the active 3- and 6-position of the pendant carbazole units yields materials with an electrochemically stable behaviour against oxidation.

Introduction

Recent developments of efficient emitters for organic light emitting diodes (OLEDs) are often focused on phosphorescent transition metal complexes. Due to elementary spin statistics 75% triplet excitons and 25% singlet excitons are formed on initial charge recombination. By fast intersystem crossing all singlet excitons will be efficiently converted to the triplet state. Hence, with these phosphorescent emitters the theoretical limit of the internal quantum efficiency is 100%.^{1–3} Due to concentration quenching effects phosphorescent materials show a loss in efficiency if the neat material is used in OLEDs. To avoid self-quenching it is necessary to dope the emitters into an appropriate host. It is essential that the triplet energy $\Delta E(T_1 - S_0)$ of the host is higher than that of the emitter in order to prevent energy back transfer from the phosphorescent guest to the host. 4,4'-Bis(9-carbazolyl)-biphenyl (**CBP**) is a widely used matrix material for phosphorescent emitters. Due to its triplet energy $\Delta E(T_1 - S_0)$ of 2.55 eV,⁴ **CBP** is a suitable matrix for green phosphorescent emitters like tris(2-phenylpyridine)iridium(III) (**Ir(ppy)**₃).^{3,5} Blue emitting materials such as the commonly used bis((4,6-difluorophenyl)-pyridinato-*N*,*C*2)picolate-iridium(III) (**FIrpic**) ($\Delta E(T_1 - S_0) = 2.62$ eV⁶) require hosts with higher

triplet energies. The key to such materials is to confine the conjugated system in the host molecules. In *N,N'*-dicarbazolyl-3,5-benzene (mCP) this is accomplished by exchanging the biphenyl group by a single benzene unit in combination with *meta* conjugation instead of *para* which leads to a triplet energy of approximately 2.90 eV.^{7,8} Another approach to enlarge the triplet energy of **CBP** based materials comprises the attachment of two methyl groups in the 2- and 2'-position of the central biphenyl which leads to 4,4'-bis(9-carbazolyl)-2,2'-dimethylbiphenyl (**CDBP**) with a triplet energy of 2.79 eV.^{4,9,10} Over the years many different carbazole based host materials have been described.^{11–13} Ma *et al.* recently described a series of non-conjugated carbazole host materials where the linkage groups between two carbazole moieties were varied. This leads to a loss of conjugation in the molecules and high triplet energies.¹⁴

Another crucial requirement for the successful operation of OLEDs is the ability of the materials to form stable amorphous films.¹⁵ This property guarantees that the emitter is uniformly diluted in the host to minimize the effect of concentration quenching. In addition, the absence of grain boundaries, which may act as trap states, makes the use of organic glasses as OLED materials advantageous.^{16,17} The glass transition temperature of materials for OLED applications is ideally above 100 °C. In general, the introduction of bulky substituents hinders packing of the molecules and leads to an amorphous behaviour of the material. On the other hand, detrimental effects on the charge carrier transport properties are observed. 9-(4-*tert*-Butylphenyl)-3,6-bis(triphenylsilyl)-carbazole (CzSi), for example, combines appropriate thermal properties with a confined conjugated system; however, the hole mobility of 5×10^{-5} cm² V⁻¹ s⁻¹ is rather low.¹⁸ By exchanging the triphenylsilyl groups by trityl

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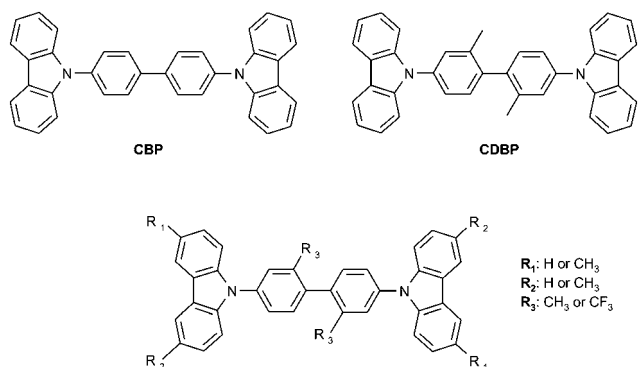
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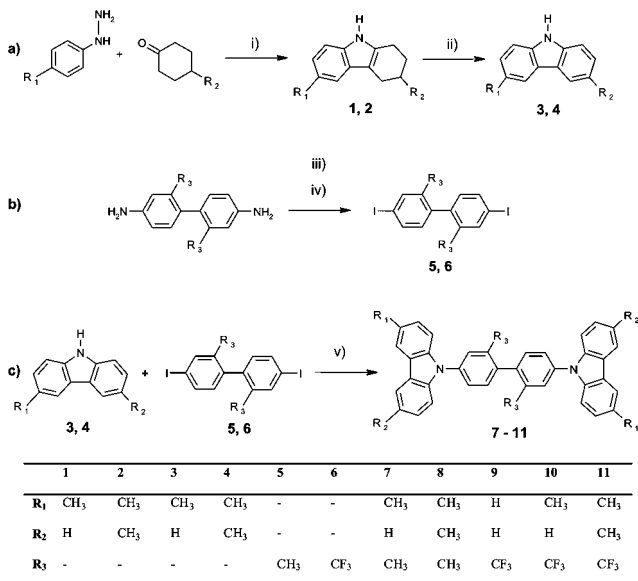
groups the hole mobility rises by one order of magnitude to $5 \times 10^{-4} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$. However, this material shows a stronger efficiency roll off at higher current densities than CzSi.¹⁹

We have prepared a series of **CBP** derivatives in which the substitution pattern in both the 2- and 2'-position of the biphenyl unit and in the 3- and 6-position of the carbazole unit has been systematically varied (Scheme 1). The attachment of the methyl and trifluoromethyl groups in these positions leads to amorphous materials with large triplet energies making them suitable as hosts for blue phosphorescent emitters.

In this work we report the synthesis of five amorphous derivatives of **CBP** together with their thermal and optical properties. The energy levels have been measured by cyclic voltammetry and absorption measurements. Furthermore, a detailed cyclic voltammetry study of the materials gives insight into the



Scheme 1 Chemical structures of the substituted 4,4'-bis(9-carbazolyl)biphenyls.



Scheme 2 Synthetic routes to (a) methyl substituted carbazoles, (b) tilted biphenyls and (c) methyl substituted **CBP** derivatives. *Reagents and conditions:* (i) acetic acid, 80 °C, 30 min; (ii) Pd(C), 1,2,4-trimethylbenzene, 170 °C, 6 h; (iii) H₂O, HCl, NaNO₂, 0–5 °C; (iv) I₂, NaI, dichloromethane, rt, 24 h; (v) Cu, K₂CO₃, 18-crown-6, *o*-dichlorobenzene, reflux, 24 h.

electrochemical stability of the **CBP** derivatives. In addition, DFT calculations were carried out to obtain the ground state geometries of the **CBP** derivatives.

Results and discussion

Synthesis

In a two step synthesis 3-methylcarbazole **3** and 3,6-dimethylcarbazole **4** were prepared from 4-methyl-phenylhydrazine and cyclohexanone or 4-methyl-cyclohexanone as the starting materials in a *Borsche* reaction (Scheme 2a). The first step yields the 1,2,3,4-tetrahydrocarbazoles **1** and **2** which are subsequently dehydrogenated with palladium on activated charcoal.^{20,21}

The synthesis of the diiodobiphenyls is shown in Scheme 2b. 4,4'-Diiodo-2,2'-dimethylbiphenyl **5** and 4,4'-diiodo-2,2'-bis(trifluoromethyl)biphenyl **6** were prepared by diazotization of the corresponding diamine and subsequent reaction with sodium iodide.²² The carbazole containing host materials **7–11** were prepared *via* the Ullmann coupling reaction of diiodo compound **5** or **6** with carbazole or the methylcarbazoles **3** and **4**, respectively (Scheme 2c).

Mass spectrometry, ¹H-NMR spectroscopy and ¹³C-NMR spectroscopy were used to identify the materials and the data are given in the Experimental section. The purity of the materials was monitored by SEC measurements.

Thermal properties

The thermal properties of the newly synthesized compounds were examined by thermogravimetric analysis (TGA) and differential scanning calorimetry (DSC) in a nitrogen atmosphere at a scanning rate of 10 K min⁻¹. All materials reveal high thermal stabilities, with an onset of weight loss at temperatures (*T*_{ID}) exceeding 310 °C, as determined by TGA measurements (Table 1).

In Fig. 1, the DSC thermograms of **CBP**, the methyl substituted compound **8** and the trifluoromethyl derivative **11** are presented. The parent compound **CBP** shows a crystalline behaviour in the DSC. The melting peak is observed at 285 °C and upon cooling the material crystallises at 183 °C. In contrast,

Table 1 Thermal properties of the carbazole-substituted biphenyls **CBP**, **CDBP**, and **7–11**^a

Entry	<i>T</i> _g ^b /°C	<i>T</i> _m ^c /°C	<i>T</i> _{cr} ^d /°C	<i>T</i> _{ID} ^a /°C
CBP	—	283	205	365
CDBP	94	—	—	310
7	106	—	—	310
8	121	277	200 ^c	312
9	100	232	—	310
10	105	210 ^b	—	337
11	119	233 ^b	—	333

^a *T*_{ID} is the temperature at which an initial loss of mass was observed in a thermogravimetric experiment with a heating rate of 10 K min⁻¹ in a nitrogen atmosphere. ^b Observed only in the first heating scan. ^c Observed during the heating scan. ^d *T*_g: glass-transition temperatures, *T*_m: melting temperatures, *T*_{cr}: crystallisation temperatures and *T*_{ID}: initial decomposition temperatures.

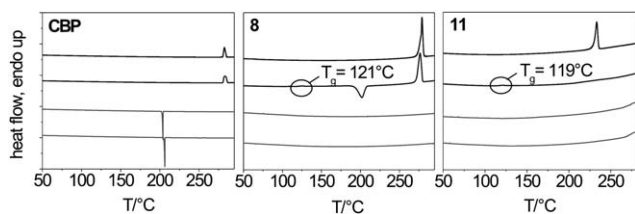


Fig. 1 DSC traces of carbazole based compounds **CBP**, **8**, and **11** at a scan rate of 10 K min⁻¹, N₂ atmosphere. Shown are from top to bottom: the first heating, second heating, first cooling and second cooling traces.

compound **8** exhibits a melting peak at 277 °C. In the cooling cycle no crystallisation is observed as the material solidifies in an amorphous phase. In the second heating curve the glass transition is observed at 121 °C followed by a recrystallisation at 200 °C. The CF₃-substituted derivative **11** remains in the amorphous phase after the first melting at 233 °C. In the second heating only the glass transition at 119 °C is observed.

These results show that by the introduction of additional methyl- and trifluoromethyl groups into the basic **CBP** structure the thermal properties are improved. While the parent compound **CBP** is highly crystalline, all **CBP** derivatives **7–11** reveal high glass transition temperatures ranging from 94 to 121 °C and all materials remain amorphous upon cooling (Table 1). For OLED applications materials with high glass transition temperatures ($T_g \approx 100$ °C) are advantageous for the operational stability of the device. In a morphologically stable amorphous host material the emitter molecules are homogeneously diluted which prevents concentration quenching. The melting, crystallisation and glass transition temperatures of all derivatives are summarised in Table 1.

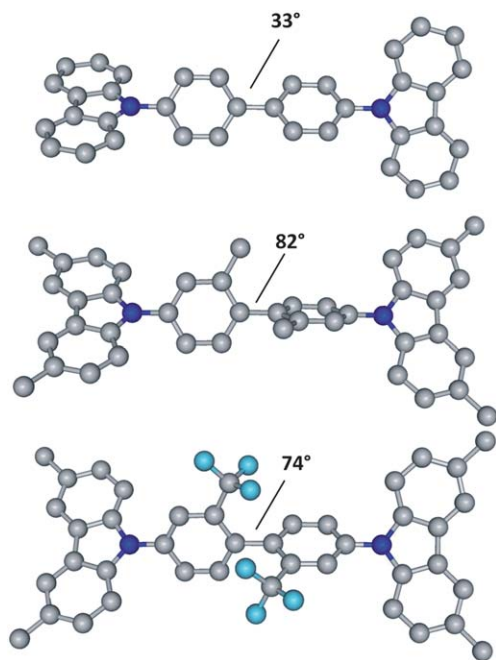


Fig. 2 Geometry optimized structures of **CBP**, **8**, and **11** (top to bottom) with different torsion angles.

Optical properties

The compounds **CBP**, **CDBP** and **7–11** were analyzed by UV/Vis, fluorescence and phosphorescence spectroscopy. In addition, the molecular structures of all compounds were simulated *via* DFT-calculations to facilitate the interpretation of the experimental results. The geometry optimized structures of **CBP**, **8**, and **11** are visualized in Fig. 2. Due to the steric demand of the substituents at the 2- and 2'-position at the biphenyl unit, **8** and **11** reveal twisted molecular structures with torsion angles between the two phenyl rings of the biphenyl unit of 82° and 74° which are significantly higher compared to the more planar **CBP** (33°).

Fig. 3 displays the room temperature absorption and fluorescence spectra taken from 10⁻⁵ M cyclohexane solutions together with the phosphorescence spectra obtained from 10 wt% solid solutions of **CBP**, **8**, and **11** in PMMA at 10 K. In order to understand how the substitutions affect the excited states of these compounds, we first consider the effect of the methyl substitution at the central biphenyl unit by comparing **CBP** and **8**. For **8**, the first and second absorption are at 351 nm and 300 nm which can both be assigned to the absorption of the 3,6-dimethylcarbazole units. In the case of **CBP**, the features associated with transitions localized on the carbazole are still present at 339 nm and 290 nm. The observed bathochromic shift of approximately 10 nm in compound **8** is caused by the additional methyl groups at the

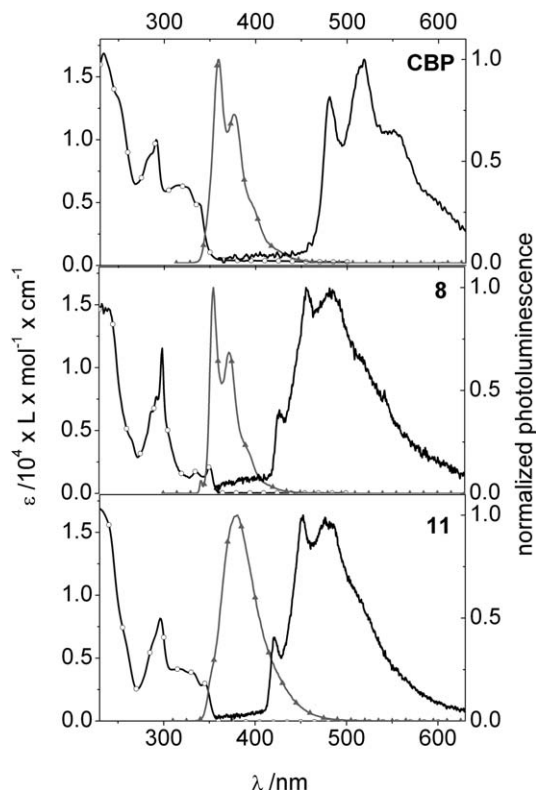


Fig. 3 Comparison of the absorption (dark grey with circles), fluorescence (light grey with triangles) and phosphorescence (black) spectra of **CBP**, **8**, and **11**. Absorption and fluorescence were taken in 10⁻⁵ M cyclohexane solutions at 300 K, phosphorescence was measured in 10 wt% solid solutions in PMMA at 10 K.

carbazole units. In **CBP**, however, there is an additional broad absorption centred at about 320 nm. This feature is likely to be associated with transitions between orbitals that involve the central biphenyl unit of the molecule. In the fluorescence spectra, in contrast, **CBP** and **8** show very similar vibrational structures with a strong maximum at about 355 nm for both compounds. We now consider how changing the substituents on the central biphenyl unit from methyl to trifluoromethyl affects the optical transitions. While **8** shows the signature of the carbazole-based transitions in the absorption and fluorescence spectra, **11** has an additional contribution to the absorption, centred at about 325 nm. As for **CBP**, we consider the additional absorption to involve the central biphenyl unit. The main difference of compound **11** compared to **CBP** and **8** is the observed featureless fluorescence centred at 410 nm. The broad bathochromically shifted fluorescence is clearly not due to transitions localized on the carbazole. The absence of vibrational fine structure rather points to a charge-transfer type transition, for example from the carbazole moiety to the central, trifluoromethyl substituted biphenyl rings.

For applications of compounds **7–11** as host materials for blue phosphorescent emitters the triplet energies are of major interest. In the phosphorescence spectra of compounds **8** and **11**, we observe two sharp peaks at 420 nm and 450 nm as well as a broad peak centred at 480 nm. On the contrary, the emission of **CBP** is bathochromically shifted by 60 nm. As a result, **8** and **11** reveal a triplet energy of approximately 2.95 eV (420 nm), significantly higher than that of **CBP** of 2.58 eV (480 nm). Apparently, the conjugation between the two phenyl rings is interrupted by introducing CH₃ and CF₃ substituents at the 2- and 2'-position of the central biphenyl leading to higher triplet energies.

All results of the photophysical investigations of compounds **CBP**, **CDBP**, and **7–11** are summarised in Table 2. With increasing CH₃-content on the pendant carbazole units in the series **CDBP**, **7**, and **8** as well as in the series of **9**, **10**, and **11** a small bathochromic shift is noticeable in both the absorption and fluorescence spectra. We attribute this to the electron donating effect of the CH₃-units. All compounds which are twisted by substituents at the central biphenyl show triplet energies of approximately 2.95 eV. These high triplet energies of the materials **CDBP** and **7–11** make them suitable host materials for deep blue phosphorescent emitters. It should be noted that the triplet energies of **CBP** and **CDBP** are slightly higher than in neat films reported by Tokito *et al.*^{4,9,10} as in our case the

phosphorescence spectroscopy was carried out in 10 wt% solid solutions of PMMA.

Cyclic voltammetry

The electrochemical behaviour of the **CBP**-derivatives was studied by cyclic voltammetry in a conventional three-electrode cell using a platinum working electrode, a platinum wire counter electrode, and a Ag/AgNO₃ reference electrode. In particular, the oxidation processes were investigated in dichloromethane solutions. The HOMO levels of the compounds were estimated from the half-wave potential of the first oxidation relative to ferrocene. The LUMO levels were calculated by adding the optical band gap to the HOMO levels. Table 3 lists the values for the HOMO and LUMO levels. In the CV-experiments **CBP** and **CDBP** show very similar HOMO levels of 5.63 eV and 5.64 eV. This observation is not surprising as in both molecules the HOMO is mainly located on the electron rich carbazole units. In the case of the methyl substitution in the 3- and 6-position of the carbazole units the additional +I-effect on the carbazole shifts the HOMO levels to slightly higher values. The HOMO level is 5.56 eV for compound **7** and 5.52 eV for compound **8**. In the case of the compounds **9–11** the strong –I-effect of the trifluoromethyl substituents causes a decrease of the HOMO level of approximately 0.1 eV (5.74 eV for **9**) in comparison with **CDBP** (5.64 eV). By the subsequent introduction of methyl groups at the carbazole moieties in compounds **10** and **11** the HOMO level rises again to 5.68 eV and 5.65 eV, respectively. In contrast, the LUMO levels of all compounds are less affected by the introduction of the substituents on the carbazole units. However, in

Table 3 Experimentally determined energy levels of the **CBP** derivatives

Entry	HOMO ^a /eV	LUMO ^b /eV
CBP	5.63	2.16
CDBP	5.64	2.13
7	5.56	2.13
8	5.52	2.13
9	5.74	2.20
10	5.68	2.22
11	5.65	2.21

^a Estimated from the half-wave potential of the first oxidation in the cyclic voltammetry measurements. ^b Estimated from the HOMO values and the optical band gap.

Table 2 Optical properties of **CBP**, **CDBP** and the **CBP**-derivatives **7–11**

Entry	λ_{EA}^a /nm, solution	λ_{EA}^b /nm, film	$\lambda_{\text{em}}^{\text{RT}c}$ /nm	$\lambda_{\text{em}}^{10\text{K}d}$ /nm	$\Delta E (S_0 - S_1)^e$ /eV	$\Delta E (T_1 - S_0)^e$ /eV
CBP	350	357	356, 374	477	3.47	2.58
CDBP	346	353	342, 358	419	3.51	2.95
7	353	361	349, 366	417	3.43	2.97
8	357	365	355, 372	421	3.39	2.95
9	343	350	368	419	3.54	2.95
10	350	358	373	419	3.46	2.95
11	353	360	379	420	3.44	2.95

^a Edge of absorption measured in 10^{−5} M cyclohexane solutions at room temperature. ^b Edge of absorption measured on neat films at room temperature. ^c Wavelengths of the intensity maxima of the fluorescence at 300 nm excitation of 10^{−5} M cyclohexane solutions at room temperature. ^d Wavelength of the highest energy maximum measured on film samples of 10 wt% compound in PMMA at 10 K. ^e The optical band gap was determined from the UV/Vis absorption onset of neat films.

compounds **9–11** the trifluoromethyl substitution at the biphenyl lowers the LUMO level slightly by approximately 8 meV compared to compounds **CBDP**, **7**, and **8**.

These considerations show that the energy levels of the **CBP** derivatives can be fine tuned to some extent by the variation of the substitution pattern at the connecting biphenyl moiety as well as at the pendant carbazoles. Especially, the HOMO levels can be varied as can be seen in Fig. 4. Thus, with these slight variations in the molecular structure the energy levels of the different layers in an OLED can be adjusted to each other in order to minimize energy barriers within the device.

Cyclic voltammetry experiments with repeated cycles give an insight into the electrochemical stability of the **CBP** derivatives. As radical cations and anions are involved in charge transport processes the electrochemical stability of the materials used in electroluminescent devices contributes to the overall stability of the device. Fig. 5 shows the cyclic voltammograms with five repeated oxidation cycles of **CBP**, **7**, and **8** in 2×10^{-3} M dichloromethane solutions.

CBP reveals an irreversible oxidation behaviour as here oligomerisation reactions of the oxidised species can take place at the active 3- and 6-position of the carbazole units. This kind of dimerisation is known from triphenylamine and *N*-phenylcarbazole. The mechanism is shown in Fig. 6.^{23,24} In the first step the molecule is oxidised at the electron lone pair of the nitrogen atom and a radical cation is formed. The radical stabilises into the 3-position of the carbazole where recombination of two radical molecules takes place under elimination of two protons.

The oligomeric species are oxidised more easily, *i.e.* at lower voltages. In the cyclic voltammogram the emerging signal at 0.6 V to 0.7 V is assigned to the oxidation of the newly formed

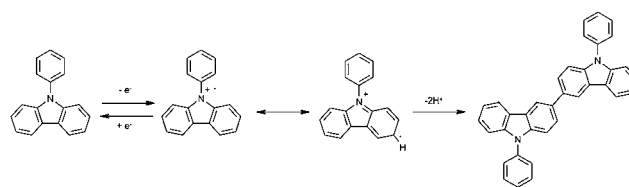


Fig. 6 Formation of the radical cation of *N*-phenylcarbazole upon oxidation and subsequent dimerisation at the active 3-position.

oligomeric species. The electrochemical behaviour of **7** is very similar to **CBP**. However, the signal between 0.4 and 0.6 V indicating the coupling of two carbazole units is growing more slowly compared with **CBP**, as in compound **7** oligomerisation can occur only at the unblocked 6-position. In contrast, compound **8** reveals a fully reversible oxidation behavior which can be attributed to the complete blocking of the active 3- and 6-position of the carbazole units.

Experimental section

Materials

All chemicals and reagents were used as received from commercial sources without further purification. 2,2'-Bis(trifluoromethyl)benzidine was synthesised according to a procedure described by Rogers *et al.*²⁵ The solvents for reactions and purification were all distilled before use.

Characterization

¹H- and ¹³C-NMR spectra were recorded with a Bruker AC 300 (300 MHz, 75 MHz) and CDCl₃ as a solvent. All data are given as chemical shifts δ (ppm) downfield from Si(CH₃)₄. For optical measurements, 10^{-5} M cyclohexane solutions of the materials as well as thin films on quartz substrates were prepared. Both neat films and films with 10 wt% of the compound in poly(methylmethacrylate) were prepared by spincoating. The UV/Vis spectra were measured in solution and on neat films with a Hitachi U-3000 spectrometer. Fluorescence spectra in solution were obtained from a Shimadzu spectrofluorophotometer RF-5301PC using excitation at 300 nm. The phosphorescence spectra were taken with the thin film samples mounted in a continuous flow helium cryostat. The temperature was controlled with an Oxford Intelligent temperature controller-4 (ITC-502). Excitation was provided by a pulsed, frequency-tripled Nd:YAG laser

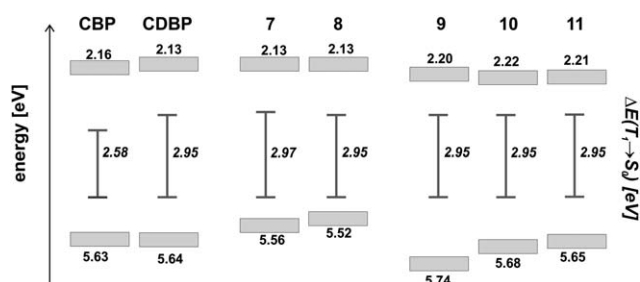


Fig. 4 Energy diagram showing the location of the HOMO and LUMO levels of the different **CBP**-derivatives. The solid line displays the position of the triplet energies $\Delta E(T_1 - S_0)$.

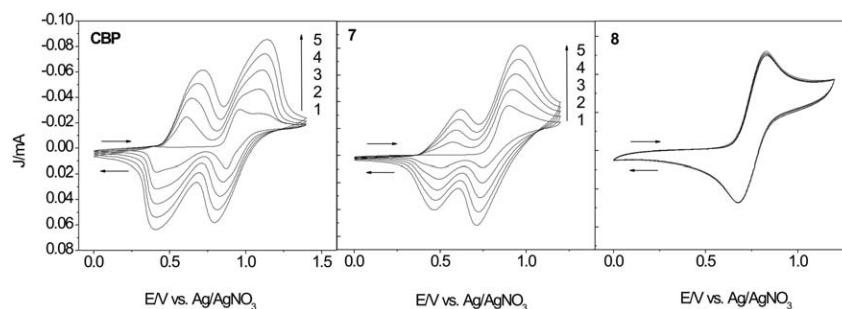


Fig. 5 Cyclic voltammograms of **CBP**, **7**, and **8** (five scans, scan rate 50 mV s⁻¹, 2×10^{-3} M in CH₂Cl₂).

at 355 nm (3.49 eV) (Spectron SL401). This wavelength corresponds to the red tail of the first absorption band in our compounds. The duration of the laser pulses was 6 ns and the laser was operated at a repetition rate of 10 Hz by a self-made electronic delay generator. The light emitted by the sample was dispersed and subsequently detected by a time gated intensified CCD camera (Andor iStar DH734-18F-9AM). The measurements were taken with a delay time of 500 ns and a gate width of 60 ms. The measurements were carried out at an excitation density of about $250 \mu\text{J cm}^{-2} \text{ pulse}^{-1}$ on films of about 150 nm thickness as determined by a Dektak profilometer. To increase the signal-to-noise-ratio, all spectra were obtained by averaging over 2000 laser shots. For differential scanning calorimetry (DSC) measurements a Diamond DSC apparatus from Perkin Elmer was used (heating/cooling rate 10 K min^{-1}). Thermogravimetric analysis (TGA) was performed on a Mettler Toledo TGA/SDTA815e machine at a heating rate of 10 K min^{-1} in a nitrogen atmosphere. Cyclic voltammetry measurements were carried out in absolute solvents measuring at a platinum working electrode versus a Ag/AgNO_3 reference electrode. Each measurement was calibrated against an internal standard (ferrocene/ferrocenium redox system). The purity of the target compounds was checked with a Waters size exclusion chromatography system (SEC) for oligomers (analytical columns: crosslinked polystyrene gel (Polymer Laboratories), length: $2 \times 60 \text{ cm}$, width: 0.8 cm , particle size: $5 \mu\text{m}$, pore size: 100 \AA , eluent: THF (0.5 mL min^{-1} , 80 bar), polystyrene standard).

Calculations

Geometries were optimized using the BP86-functional^{26,27} in combination with a split-valence basis set (SV(P)) including polarization functions on all heavy atoms.²⁸ All calculations were carried out with the turbomole program package.²⁹

Synthetic procedures

General procedure for the preparation of methyl substituted carbazoles

3-Methyl-6,7,8,9-tetrahydro-5H-carbazole (1). 3-Methylphenylhydrazine hydrochloride (10.0 g, 63 mmol) was added to a solution of cyclohexanone (6.2 g, 63 mmol) in acetic acid (30 mL) under an argon atmosphere over a period of one hour. After stirring at 80°C for half an hour the reaction mixture was extracted with dichloromethane and washed several times with 5% aqueous sodium hydrogen carbonate solution to neutralise the acetic acid. After drying the organic layer over anhydrous sodium sulfate the solvent was evaporated. Yield: 9.25 g (79%). EI-MS m/z : 185 (95, M^+).

3,6-Dimethyl-6,7,8,9-tetrahydro-5H-carbazole (2). Compound **2** was prepared according to the procedure described for **1**. Yield: 93%.

3-Methylcarbazole (3). To 9.25 g (50 mmol) of 3-methyl-6,7,8,9-tetrahydro-5H-carbazole (**1**) in 30 mL of 1,2,4-trimethylbenzene was added 10% palladium on activated charcoal (2.66 g, 25 mmol). The mixture was refluxed for 6 h. In order to remove the catalyst the mixture was diluted with

dichloromethane and filtered over neutral aluminium oxide. After the removal of dichloromethane, hexane was added and the product was obtained as white precipitate. Yield: 88%. EI-MS m/z : 181 (100, M^+). $^1\text{H-NMR}$ (300 MHz, CDCl_3), δ (ppm): 8.06 (dd, 1H), 7.94 (s, 1H, NH), 7.90 (ds, 1H), 7.41 (m, 2H), 7.35 (d, 1H), 7.27–7.21 (m, 2H), 2.55 (s, 3H). $^{13}\text{C-NMR}$ (75 MHz, CDCl_3), δ (ppm): 139.82, 137.71, 137.26, 128.75, 127.18, 125.65, 123.14, 120.26, 120.24, 119.22, 110.55, 110.24, 21.44.

3,6-Dimethylcarbazole (4). Compound **4** was prepared according to the procedure given for **3**. Yield: 86%. EI-MS m/z : 195 (100, M^+). $^1\text{H-NMR}$ (300 MHz, CDCl_3), δ (ppm): 7.85 (m, 2H), 7.84 (s, 1H, NH), 7.31 (d, 2H), 7.24 (dd, 2H), 2.55 (s, 6H). $^{13}\text{C-NMR}$ (75 MHz, CDCl_3), δ (ppm): 138.07, 128.48, 126.99, 123.41, 120.18, 110.22, 21.44.

General procedure for the preparation of substituted biphenyls

4,4'-Diiodo-2,2'-dimethylbiphenyl (5). 2,2'-Dimethylbenzidine dihydrochloride (14.3 g, 50 mmol) was suspended in 125 mL of water and 15 mL of conc. hydrochloric acid and the suspension was cooled to $0-5^\circ\text{C}$. A solution of sodium nitrite (7.18 g, 100 mmol) in 20 mL of water was added dropwise. The resulting cold tetrazonium salt solution was added slowly to a well-stirred solution of iodine (30.5 g, 120 mmol) and sodium iodide (30.0 g, 200 mmol) in 50 mL of water and 100 mL of dichloromethane at a temperature below 5°C . The mixture was stirred for 24 h at room temperature and the excess of iodine was removed by the addition of a sodium thiosulfate solution. The product was extracted with dichloromethane and washed several times with water. After drying the organic layer over anhydrous sodium sulfate, the solvent was evaporated and the residue was purified by column chromatography on silica gel with hexane as eluent to afford compound **5**. Yield: 16.6 g (76%). EI-MS m/z : 434 (100, M^+). $^1\text{H-NMR}$ (300 MHz, CDCl_3), δ (ppm): 7.63 (d, 2H), 7.55 (dd, 2H), 6.79 (d, 2H), 1.98 (s, 6H). $^{13}\text{C-NMR}$ (75 MHz, CDCl_3), δ (ppm): 139.99, 138.75, 138.22, 134.81, 130.87, 93.13, 19.47.

4,4'-Diiodo-2,2'-bis(trifluoromethyl)biphenyl (6). Compound **6** was prepared according to the procedure given for **5**. Yield: 1.91 g (75%). EI-MS m/z : 542 (100, M^+). $^1\text{H-NMR}$ (300 MHz, CDCl_3), δ (ppm): 8.07 (d, 2H), 7.89 (dd, 2H), 6.99 (d, 2H). $^{13}\text{C-NMR}$ (75 MHz, CDCl_3), δ (ppm): 139.97, 135.88, 135.06, 132.86, 130.32 (q, $^2J(\text{C-F}) = 30.8 \text{ Hz}$), 122.56 (q, $^1J(\text{C-F}) = 273.0 \text{ Hz}$), 93.55.

General procedure for the Ullmann condensation

4,4'-Bis(3-methylcarbazolyl)-2,2'-dimethylbiphenyl (7). 4,4'-Diiodo-2,2'-dimethylbiphenyl (**5**) (1 g, 2.3 mmol), 3-methylcarbazole (**3**) (1.0 g, 5.52 mmol), potassium carbonate (2.5 g, 18.4 mmol), copper powder (0.58 g, 9.2 mmol) and 18-crown-6 (0.12 mg, 0.46 mmol) were refluxed in 15 mL of *o*-dichlorobenzene in an argon atmosphere for 24 h. Copper and inorganic salts were filtered off and the solvent was evaporated. Column chromatography on silica gel with hexane/tetrahydrofuran (20 : 1) as eluent yielded 0.72 g (58%) of **7** as white solid. EI-MS m/z : 540 (100, M^+). $^1\text{H-NMR}$ (300 MHz, CDCl_3), δ (ppm): 8.14 (d, 4H), 7.97 (m, 4H), 7.55–7.43 (m, 12H), 7.32–7.27 (m, 4H), 2.58 (s, 6H), 2.29 (s, 6H). $^{13}\text{C-NMR}$ (75 MHz, CDCl_3), δ (ppm): 141.11, 139.83, 139.23, 137.88, 137.11, 130.78, 129.31, 128.19, 127.24,

125.76, 124.16, 123.57, 123.30, 120.28, 119.72, 109.88, 109.65, 21.45, 20.17. T_m : $-^{\circ}\text{C}$; T_g : 106°C .

4,4'-Bis(9-carbazolyl)-2,2'-dimethylbiphenyl (**CDBP**), 4,4'-bis(3,6-dimethylcarbazolyl)-2,2'-dimethylbiphenyl (**8**), 4,4'-bis(9-carbazolyl)-2,2'-bis(trifluoromethyl)biphenyl (**9**),³⁰ 4,4'-bis(3-methylcarbazolyl)-2,2'-bis(trifluoromethyl)biphenyl (**10**), and 4,4'-bis(3,6-dimethylcarbazolyl)-2,2'-bis(trifluoromethyl)biphenyl (**11**) were prepared according to the procedure given for **7**.

4,4'-Bis(9-carbazolyl)-2,2'-dimethylbiphenyl (**CDBP**). Yield: 0.59 g (50%). EI-MS m/z : 512 (100, M^+). $^1\text{H-NMR}$ (300 MHz, CDCl_3), δ (ppm): 8.20 (d, 4H), 7.58–7.45 (m, 14H), 7.36–7.31 (m, 4H), 2.31 (s, 6H). $^{13}\text{C-NMR}$ (75 MHz, CDCl_3), δ (ppm): 140.96, 140.01, 137.95, 136.92, 130.81, 128.36, 125.94, 124.34, 123.42, 120.36, 119.94, 109.94, 20.17. T_m : $-^{\circ}\text{C}$; T_g : 94°C .

4,4'-Bis(3,6-dimethylcarbazolyl)-2,2'-dimethylbiphenyl (**8**). Yield: 0.7 g (50%). EI-MS m/z : 568 (100, M^+). $^1\text{H-NMR}$ (300 MHz, CDCl_3), δ (ppm): 7.93 (ds, 4H), 7.52–7.40 (m, 10H), 7.26 (dd, 4H), 2.57 (s, 12H), 2.28 (s, 6H). $^{13}\text{C-NMR}$ (75 MHz, CDCl_3), δ (ppm): 139.64, 139.40, 137.82, 137.32, 130.75, 129.06, 128.01, 127.07, 123.98, 123.46, 120.22, 109.60, 21.45, 20.19. T_m : 277°C ; T_g : 121°C .

4,4'-Bis(9-carbazolyl)-2,2'-bis(trifluoromethyl)biphenyl (**9**). Yield: 0.71 g (62%). EI-MS m/z : 620 (100, M^+). $^1\text{H-NMR}$ (300 MHz, CDCl_3), δ (ppm): 8.20 (d, 4H), 8.09 (d, 2H), 7.93 (dd, 2H), 7.71 (d, 2H), 7.55–7.48 (m, 8H), 7.40–7.35 (m, 4H). $^{13}\text{C-NMR}$ (75 MHz, CDCl_3), δ (ppm): 140.41, 138.33, 135.32, 133.49, 133.41, 130.90 (q, $^2J(\text{C-F}) = 31.5$ Hz), 129.15, 126.48, 124.68, 123.83, 123.44 (q, $^1J(\text{C-F}) = 272.3$ Hz), 120.76, 120.60, 109.50. T_m : 232°C ; T_g : 100°C .

4,4'-Bis(3-methylcarbazolyl)-2,2'-bis(trifluoromethyl)biphenyl (**10**). Yield: 0.8 g (55%). EI-MS m/z : 648 (100, M^+). $^1\text{H-NMR}$ (300 MHz, CDCl_3), δ (ppm): 8.15 (d, 2H), 8.07 (d, 2H), 7.99 (m, 2H), 7.86 (dd, 2H), 7.68 (d, 2H), 7.54–7.30 (m, 10H), 2.60 (s, 6H). $^{13}\text{C-NMR}$ (75 MHz, CDCl_3), δ (ppm): 140.56, 138.66, 138.51, 135.10, 133.37, 130.82 (q, $^2J(\text{C-F}) = 30.8$ Hz), 130.22, 128.90, 127.63, 126.20, 124.48, 123.99, 123.74, 123.46 (q, $^1J(\text{C-F}) = 273.8$ Hz), 120.56, 120.50, 109.45, 109.22, 21.43. T_m : 210°C ; T_g : 105°C .

4,4'-Bis(3,6-dimethylcarbazolyl)-2,2'-bis(trifluoromethyl)biphenyl (**11**). Yield: 1.4 g (56%). EI-MS m/z : 676 (100, M^+). $^1\text{H-NMR}$ (300 MHz, CDCl_3), δ (ppm): 8.05 (d, 2H), 7.95 (m, 4H), 7.85 (dd, 2H), 7.66 (d, 2H), 7.42 (d, 4H), 7.30 (dd, 4H), 2.59 (s, 12H). $^{13}\text{C-NMR}$ (75 MHz, CDCl_3), δ (ppm): 138.82, 138.71, 134.87, 133.34, 130.75 (q, $^2J(\text{C-F}) = 30.8$ Hz), 130.00, 128.66, 127.46, 124.25, 123.91, 123.50 (q, $^1J(\text{C-F}) = 272.3$ Hz), 120.46, 109.18, 21.43. T_m : 233°C ; T_g : 119°C .

Conclusions

We have described a number of **CBP** derivatives in which the substitution in the 2- and 2'-position of the biphenyl and in the 3- and 6-position of the carbazole unit is systematically varied. This substitution of the parent crystalline **CBP** leads to the materials

7–11 which have a much lower tendency to crystallise. Their glass transition temperatures range from 94 – 121°C . The highest glass transition temperatures were determined for 3,6-dimethylcarbazole bearing derivatives **8** and **11** with 121°C and 119°C , respectively.

The main effect of the methyl substitution in 2- and 2'-position of the biphenyl unit is a twisting of the two central phenyl rings. Due to this electronic decoupling, the conjugation length in the molecule is limited which causes an increase of the triplet energy $\Delta E(T_1 - S_0)$ from 2.58 eV for **CBP** to 2.95 – 2.97 eV for **7–11**.

Fine-tuning of the energy levels, especially the HOMO levels, can be achieved by a suitable choice of substitution pattern of the **CBP** derivatives. Cyclic voltammetry with repeated cycles shows that by introducing substituents at the 3- and 6-position of the pendant carbazole units the oxidation to the radical cation becomes fully reversible and thus electrochemical stable host materials are accessible.

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